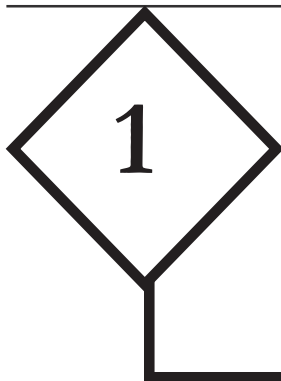


INDEX

Chemistry

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9. **OILS AND FATS**
10. **CHEMISTRY & INDUSTRY**



UNIT - I

Atomic Structure

MULTIPLE CHOICE QUESTIONS

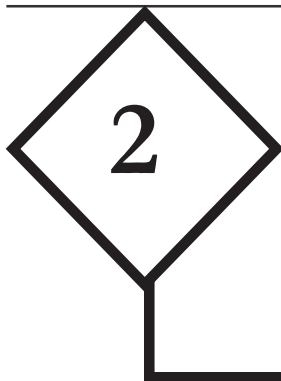
- Bohr's theory is valid for
 - all atoms
 - all ions
 - any atom or ion having 1 electron
 - all molecules
- Which of the following shells has the least energy?
 - L
 - M
 - K
 - N
- Who introduced elliptical orbits?
 - Bohr
 - Schrodinger
 - Zeeman
 - Sommerfeld
- Which of the following quantum numbers give size and energy of stationary orbit?
 - n
 - l
 - m
 - s
- The sub-shells present in L-shell are
 - s and d
 - s and p
 - s, p and d
 - s, p, d and f
- The maximum value of l for n=5 is
 - 5
 - 3
 - 5
 - 4
- The number of d-orbitals present in n=3 is
 - 1
 - 3
 - 5
 - 7
- f-orbitals are present in
 - K-shell
 - L-shell
 - M-shell
 - N-shell
- Among 3p, 4s, 3d and 4p, the orbital having least energy is
 - 4s
 - 3p
 - 3d
 - 4p
- The 'l' value of 'd' sub-shell is
 - 0
 - 1
 - 2
 - 3
- The clockwise spin of electron is represented by
 -
 -
 -
 -
- Shape of s-orbital is
 - dumbbell
 - spherical
 - double dumb bell
 - tetrahedral
- Elliptical orbits were introduced by
 - Bohr
 - Schrodinger
 - Zeeman
 - Sommerfeld
- The number of 'm' values for l=3 is
 - 7
 - 4
 - 2
 - 3
- Among sub atomic particles which are known as nucleons
 - Protons + Electrons
 - Protons + Neutrons
 - Neutrons+ Electrons
 - Electrons + Positrons

16. Electron pairing takes place only after all the available degenerate orbitals are occupied by one electron each. This is known as
 a) Aufbau principle b) Pauli's exclusion principle c) Hund's rule d) Avogadro's hypothesis
17. Rutherford suggested that atoms are
 a) spherical b) cylindrical c) tetrahedral d) pyramidal
18. Orbitals will have
 a) nodal regions b) antinodal regions c) crest regions d) compression regions
19. The force that pulls away the electron from the nucleus is
 a) centripetal force b) attraction between electrons and the nucleus
 c) electrostatic force d) centrifugal force
20. The value of Planck's constant is
 a) b) c) d) 6.625×10^{-34} J.sec
21. The shape of p-orbital is
 a) double-dumbbell b) spherical c) cylindrical d) dumbbell
22. The angular momentum of the electron revolving in a stationary orbit is equal to integral multiples of
 a) $2\pi / h$ b) c) d)
23. The shape of s-orbital is
 a) spherical b) elliptical c) dumbbell d) double dumbbell
24. Bohr's model could not account for
 a) seebeck effect b) raman's effect c) condensation effect d) zeeman effect
25. The region in space where there is a finite probability of finding an electron is called
 a) Atomic number b) Atomic orbital c) Atomic particle d) Ground state
26. The sub-states in a stationary state are called ~~Atomic orbitals~~
 a) orbits b) atomic orbitals c) shells d) atomic particles
27. The orbital whose l value is '0' is designated as
 a) s orbital b) p orbital c) d orbital d) f orbital
28. Orbitals having same energy are called
 a) valence shells b) empty shells c) degenerate shells d) overlappign shells
29. The quantum number that gives the information regarding the shape of substationary state is
 a) Principal quantum number b) Azimuthal quantum number
 c) Magnetic quantum number d) Spin quantum number
30. The number of states present in the sub stationary state 'd' is
 a) 1 b) 3 c) 5 d) 7
31. The number of states present in the sub stationary state of 'g' is
 a) 3 b) 5 c) 7 d) 9
32. The electronic configuration of copper is
 a) b) c) d)
33. The negatively charged particle in the atom is
 a) Electron b) Neutron c) Proton d) Positron
34. The number of sub atomic particles in an atom are
 a) 1 b) 2 c) 3 d) 100
35. Neutral particle int eh nucleus of an atom is
 a) Electron b) Neutron c) Proton d) Positron

36. The planetary model of the atom is proposed by
 a) Thomson b) Rutherford c) Bohr d) Sommerfeld
37. By removing or adding an electron to an atom it becomes a
 a) Ion b) Radical c) Molecule d) Neutron
38. Magnetic quantum number is related to
 a) Size b) Shape c) Orientation d) Spin
39. Number of electrons that can be accommodated in f sub-shell is
 a) 2 b) 8 c) 32 d) 14
40. The phenomenon of black body radiation was successfully explained by
 a) Rutherford b) Thomson c) Max planck d) Schrodinger
41. Orbital without any directional character
 a) 4 b) p c) d d) f
42. According to Sommerfeld's elliptical orbits the most penetrating orbit towards the nucleus is
 a) s b) p c) d d) f
43. Any two electrons in an atom can have
 a) equal values for n,l, m but not s b) equal values for l, m and s but not n
 c) equal values for n, m and s but not l d) all the them are true
44. The fine structure of atomic spectrum is satisfactorily explained by
 a) Bohr's atomic model b) Sommerfeld's model
 c) Wave mechanical model of an atom d) Rutherford's model of an atom
45. The application of quantum mechanics to atomic structure is based largely on the research work os
 a) Einstein b) Schrodinger c) Bohr d) Rutherford
46. "No two electrons will have all the four quantum number same". This is known as
 a) Aufbau principle b) Hund's rule c) Pauli's exclusion principle d) Ampere's rule
47. Neutrons can be found in atoms of all the elements except in
 a) Carbon b) Helium c) Hydrogen d) Oxygen
48. Maximum number of electrons in a 3d orbital is
 a) 2 b) 10 c) 6 d) 14
49. Which of the following orbitals does not exist?
 a) 6s b) 2d c) 5p d) 1s
50. For ns orbital, value of magnetic quantum number is
 a) -1 b) 0 c) +1 d) n

KEY

- | | | | | | | | |
|-------|-------|-------|-------|-------|-------|-------|-------|
| 1) c | 2) c | 3) d | 4) a | 5) b | 6) d | 7) c | 8) d |
| 9) b | 10) c | 11) d | 12) b | 13) b | 14) a | 15) b | 16) c |
| 17) a | 18) a | 19) d | 20) d | 21) d | 22) b | 23) a | 24) d |
| 25) b | 26) b | 27) a | 28) c | 29) b | 30) c | 31) d | 32) b |
| 33) a | 34) c | 35) b | 36) b | 37) a | 38) c | 39) d | 40) c |
| 41) a | 42) a | 43) d | 44) c | 45) b | 46) c | 47) c | 48) b |
| 49) b | 50) b | | | | | | |



UNIT - II

Chemical Bond

- Complete transfer of electrons from one atom to another leads to the formation of
 - ionic bond
 - covalent bond
 - coordinate covalent bond
 - none
- Co-ordinate covalent bond is present in
 - HCl
 -
 -
 -
- Shape of CH_4 is
 - 'V' shape
 - pyramidal
 - Linear
 - Tetrahedral
- s-p overlap is present in
 -
 -
 -
 -
- "V" shaped molecule is
 -
 -
 -
 -
- Bond angle in water molecule is
 -
 -
 -
 -
- s-s overlap is found in
 -
 -
 -
 -
- p-p overlap is observed in
 -
 -
 -
 - HBr
- End-on-end overlap results in
 - Pi () bond
 - sigma () bond
 - hydrogen bond
 - ionic bond
- The number of valency electrons in carbon atom ($Z=6$) is
 - 0
 - 2
 - 4
 - 6
- An element has the electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^2$. Its valence electrons are
 - 2
 - 6
 - 3
 - 4
- The octet rule is not valid for the molecule
 -
 -
 -
 -
- The valency shell of calcium contains
 - 2 electrons
 - 4 electrons
 - 6 electrons
 - 8 electrons
- Side-on overlap results in
 - sigma ()
 - Pi () bond
 - bond
 - ionic bond

15. A molecule having double bonds
 a) b) c) d)
16. Example of a molecule having Pi () bond
 a) b) c) d)
17. Which of the following is non-linear?
 a) b) c) d)
18. In a double bond
 a) Two pi () bonds b) two sigma () bonds
 c) Three sigma () bonds d) none
19. In a triple bond
 a) one sigma () and two pi () bonds b) one pi () and two sigma () bonds
 c) three sigma () bonds d) three pi () bonds
20. Which of the following has sigma bond?
 a) b) c) d)
21. Which of the following has triple bond?
 a) b) c) d)
22. Which of the following has only sigma () bond?
 a) b) c) q d) none
23. The donor in Ammonia boron tri fluoride is
 a) N b) B ~~(c) N and B~~ d) None
24. The shape of Ammonia molecule is
 a) pyramidal b) tetrahedral c) trigonal planar d) trigonal bi pyramidal
25. The shape of Boron trifluoride molecule is
 a) Planar triangular b) Linear c) Trigonal bi-pyramidal d) Tetrahedral

KEY

- 1) a 2) c 3) c 4) d 5) d 6) a 7) c 8) a
 9) b 10) c 11) d 12) c 13) a 14) b 15) a 16) b
 17) a 18) c 19) a 20) b 21) c 22) c 23) a 24) a
 25) a

3

UNIT - III

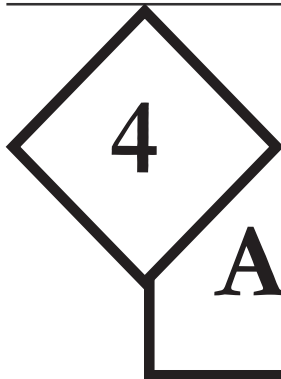
Periodic Classification

- In a Dobereiner triad, the atomic weight of middle element is
 - the sum of the atomic weights of first and the third elements
 - the product of the atomic weights of first and the third elements
 - the ratio of the atomic weights of first and the third elements
 - the mean of the atomic weight of first and the third elements
- Mendeleef's periodic table is based on the
 - atomic weight
 - atomic number
 - atomic radius
 - atomic volume
- The units of atomic radius are
 - angstroms
 - $\text{kJ}\cdot\text{mol}^{-1}$
 - eV
 - $\text{K}\cdot\text{Cal}\cdot\text{mol}^{-1}$
- The ionization potential in a group from top to bottom
 - decreases
 - increases
 - remains the same
 - increases and decreases
- The law of octaves applies to
 - B,C,N
 - As, K, Ca
 - Be, Mg, Ca
 - None
- The elements which is cited as an example to prove the validity of mendeleef's periodic law is
 - indium
 - hofnium
 - gallium
 - helium
- Which of the following have the minimum atomic radius?
 - N
 - Na
 - K
 - F
- f-block elements are also called
 - transition elemtns
 - Transuranic elements
 - Alkali metals
 - Inner transition elements
- Alkali metals are powerful
 - oxidants
 - reductants
 - radioactive elements
 - none
- In the periodic table, elements of the same group have the same
 - atomic number
 - atomic weight
 - valence electrons
 - atomic size
- Addition of oxygen to a compound is termed as
 - oxidation
 - neutralisation
 - reduction
 - dsproportionation
- The number of electrons in the outermost shell of an inert gas is
 - 4
 - 6
 - 8
 - 10
- The d-block of elements are also called
 - Representative elements
 - Transition elements
 - Inner transition elements
 - Halogens
- The lightest metal is
 - Li
 - H
 - He
 - Na

15. Size of chloride ion is
- a) bigger than chlorine atom
 - b) smaller than chlorine atom
 - c) same as chlorine atom
 - d) none

KEY

- 1) d 2) a 3) a 4) a 5) c 6) c 7) d 8) d
9) b 10) c 11) a 12) c 13) b 14) a 15) a



UNIT - IV

Alkaline Earth Metals

- _____ is the electronic configuration of
 - Be
 - Mg
 - Ca
 - Sr
- Which of the following property increases from Be to Ra?
 - EN
 - Ionization energy
 - Atomic size
 - None
- CaO is
 - acidic
 - basic
 - neutral
 - amphoteric
- Which of the following metal gives peroxide in addition to oxide when burnt in excess of air?
 - Be
 - Mg
 - Ca
 - Ba
- During the electrolytic extraction of Mg, the cathode used is
 - iron pot
 - graphite
 - KCl and NaCl
 - porcelain tube
- Which of the following is an ore of Mg?
 - Beryl
 - Barytes
 - ~~MgO~~ Mg_2O
 - Hematite
- Dolomite is a mineral of
 - Mg
 - Ca
 - Ba
 - Al
- Which is the Alkaline Earth metal in the following
 - Potassium
 - Sodium
 - Rubidium
 - Radium
- Carnallite is
 -
 -
 -
 -
- Chemical formula of Epsom salt
 -
 -
 -
 -
- The atomic size in group IIA elements from Be to Ra
 - increases
 - decreases
 - doesn't change
 - constant
- Melting point of Mg
 -
 -
 - doesn't change
 - constant
- Magnesium reacts with hot water and liberates the following gas
 -
 -
 -
 -
- Magnesium burns in air and forms the following compound in addition to MgO
 -
 -
 -
 -
- The following alkaline earth metal is used in making deepavali crackers
 - Be
 - Ca
 - Mg
 - Ba

16. Oxides of alkaline earth metals are
a) basic b) acidic c) amphoteric d) neutral
17. By reducing with lithium aluminium hydride, the following compound is obtained
a) b) c) d)
18. The chemical formula of lithium aluminium hydride is
a) b) c) d)
19. A compound which is hygroscopic
a) MgO b) c) d)
20. The following metal is obtained by electrolytic reduction of molten magnesite
a) Ca b) Sr c) Ba d) Mg

KEY

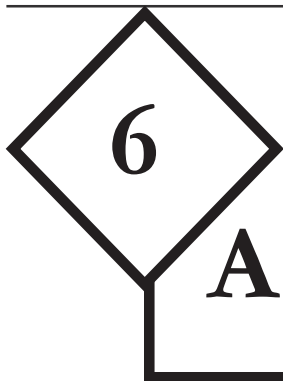
- 1) b 2) c 3) b 4) d 5) a 6) c 7) a 8) d
9) b 10) d 11) a 12) c 13) d 14) a 15) c 16) a
17) a 18) b 19) c 20) d



16. If the mole fraction of solute is 0.3, mole fraction of solvent is
 a) 0 b) 0.3 c) 0.7 d) 9.7
17. In is the mole fraction of solute and is the mole fraction of solvent, then =
 a) 0 b) 1 c) 10 d)
18. The positively charged ions when NaCl gets ionized is
 a) b) c) d)
19. The process of a molecule giving rise to ions is called
 a) reduction b) neutralisation c) oxidation d) ionization
20. An example of a strong electrolyte
 a) b) sucrose c) d)
21. One of the following is a weak electrolyte
 a) NaCl b) KCl c) d)
22. Identify the non-electrolyte among the following
 a) b) c) Urea d)
23. The extent of ionization increases by increase of
 a) light b) dilution c) concentration d) none
24. The weight of required to make 250 ml of 0.1 M solution is
 a) 1.32 gm b) 1.989 gm c) 2.65 gm d) 3.96 gm
25. When 2.3 gm of (m.wt= 46) is dissolved in 500 c.c of water, the molarity of the solution is
 a) 0.01 b) 0.05 c) 0.1 d) 2
26. A homogeneous mixture of two or more substances is called
 a) solute b) solution c) solvent d) liquor
27. How many moles of water are present in 180g of water?
 a) 1 mole b) 18 moles c) 10 moles d) 10 moles
28. A hamogeneous mixture of two or more substances is called
 a) solute b) solution c) solvent d) liquor
29. The mole fraction of A in a solution containing 0.2 mole each of A,B and C is
 a) 0.2 b) 0.333 c) 0.5 d) 0.66
30. If more solute is present in solution than required for saturation, it is called
 a) super saturated solution b) saturated solution
 c) unsaturated solution d) none

KEY

- | | | | | | | | |
|-------|-------|-------|-------|-------|-------|-------|-------|
| 1) c | 2) b | 3) c | 4) a | 5) d | 6) a | 7) b | 8) d |
| 9) b | 10) a | 11) d | 12)c | 13) d | 14) a | 15) d | 16) c |
| 17) b | 18) a | 19) d | 20) d | 21) d | 22) c | 23) b | 24) c |
| 25) c | 26) b | 27) c | 28) b | 29) b | 30) a | | |



UNIT - VI

Acids, Bases and Salts

- The colour of methyl orange indicator in acidic medium is
a) yellow b) green c) orange d) red
- The colour of phenolphthalein indicator in basic solution is
a) yellow b) green c) pink d) orange
- The p^H of solution whose p^H is 4 is
a) 1 b) 10 c) -4 d) +4
- If the p^H of solution is 8, its pOH is
a) 1 b) 10 c) 8 d) 0
- The value of p^H changes with changing
a) temperature b) pressure c) concentration d) volume
- Weak acids ionize upto
a) 50% b) 100% c) less than 100% d) more than 100%
- The heat of neutralisation for a strong acid and a strong base is
a) 13.7 k. cal b) 13.4 k. cal c) 0 k. cal d) 0.3 k. cal
- CaO is
a) acidic b) basic c) neutral d) amphoteric
- p^H of water is
a) 14 b) 7 c) 0 d) none
- If the p^H of a solution is 10, its pOH is
a) 10 b) 4 c) 0 d) 14
- Quantity of heat of neutralisation between a strong acid and a strong base is
a) 13.7 k. cal/mole b) 1.37 k. cal/mole c) 13.7 cal/mole d) 1370 cal/mole
- The p^H of a solution is defined by the expression
a) $-\log[H^+]$ b) $-\log[OH^-]$ c) $-\log[H^+][OH^-]$ d) $-\log[H^+][OH^-]$



UNIT - VII

Chemistry of Carbon Compounds

- The refractive index of diamond is
a) 4.3 b) 2.45 c) 4.5 d) 5.42
- Bond length (in units) in graphite is
a) 2.45 b) 1.42 c) 4.21 d) 2.81
- _____ gas is cooled to form dry ice by sudden expansion. This is called
a) einstein effect b) rutherford effect c) joule-thomson effect d) bohr effect
- The name of _____ is
a) hexane b) octane c) methane d) propane
- Alkanes undergo
a) addition reactions b) substitution reactions c) condensation reactions d) polymerization reactions
- COOR is called _____
a) acid group b) amine group c) ester group d) ketone group
- Which of the following is an alkane?
a) _____ b) _____ c) _____ d) _____
- An example for C-COOR functional
a) _____ b) _____ c) _____ d) _____
- The functional group of Ketone is
a) -OH b) -CHO c) -O- d) >C=O
- _____ is called
a) acid group b) amine group c) ester group d) ketone group
- Chief component of cooking gas is
a) butane b) methane c) ethane d) octane
- Dry ice is
a) solid b) solid c) solid d) solid
- Diamond and graphite are
a) isomers b) isomorphous c) isotones d) allotropes
- Hardest known material in nature is
a) diamond b) graphite c) granite d) none

15. Diamond and graphite have structures
a) tetrahedral and square planar b) tetrahedral and octahedral
c) tetrahedral and hexagonal d) hexagonal and square planar

KEY

- 1) b 2) b 3) c 4) b 5) b 6) c 7) a 8) a
9) d 10) b 11) a 12) c 13) d 14) a 15) c



UNIT - VIII

Carbohydrates and Proteins

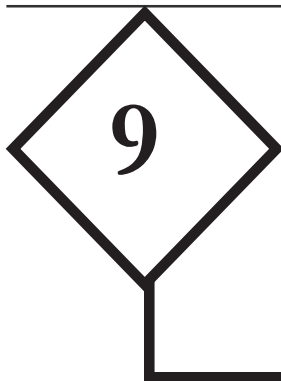
- Which of the following is the sweetest
a) sucrose b) glucose c) fructose d) maltose
- A polysaccharide is
a) Glucose b) Fructose c) Sucrose d) Starch
- Aldoses are
a) polyhydroxy ketones b) polyhydroxy aldehydes
c) polyhydroxy amines d) polyhydroxy esters
- Hexoses contain
a) 3 carbons b) 4 carbons c) 5 carbons d) 6 carbons
- In the Tollen's test glucose reduces
a) Ag metal to Ag^+ ion b) Ag^+ ion to Ag metal c) Cu^{2+} ion to Cu^+ ion d) Cu^+ ion to Cu^{2+} ion
- Defecation is addition of H_2O
a) H_2O b) H_2 c) O_2 d) CO_2
- Acidity in the juice is removed by adding
a) $NaOH$ b) $Ca(OH)_2$ c) Na_2CO_3 d) $NaHCO_3$
- The sugar content of molasses is
a) 10% b) 20% c) 50% d) 90%
- Which of the following is not a byproduct of sugar industry?
a) Bagasse b) Press mud c) Sugar d) Molasses
- The chief use of ethyl alcohol is
a) For drinking b) As solvent c) As medicine d) For making beverages
- Which of the following is used to get absolute alcohol from rectified spirit?
a) CaO b) $CaCl_2$ c) CaH_2 d) pyridine
- Consumption of denatured spirit causes
a) unconsciousness b) blindness c) ulcers in the intestine d) damage to lungs
- Benedict's solution contains
a) silver nitrate b) copper carbonate c) copper oxide d) copper sulphate
- The spent cane is called
a) Yeast b) Zymase c) bagasee d) press mud
- Sucrose is an example of
a) monosaccharides b) polysaccharides c) oligosaccharides d) none

16. Starch is an example of
a) monosaccharide b) polysaccharide c) oligosaccharide d) none
17. The general formula of polysaccharides
a) b) c) $(C_6H_{10}O_5)_n$ d) none
18. Sweetest sugar is
a) glucose b) sucrose c) lactose d) fructose
19. Yeast produces enzymes namely
a) aldose b) ketose c) invertase and zymase d) None
20. The widely used solvent next to water is
a) Sugar b) Benzene c) d) Alcohol

KEY

- 1) c 2) d 3) b 4) d 5) b 6) b 7) b 8) c
9) c 10) b 11) b 12) b 13) a 14) c 15) c 16) b
17) c 18) d 19) c 20) d





UNIT - IX

Oils and Fats

- The chief sources of oils is
a) petroleum b) coal and coke c) animals and plants d) soaps and detergents
- Which of the following is a saturated fatty acid?
a) myristoleic acid b) lauric acid c) palmitoleic acid d) linoleic acid
- The cation of soap useful for dry cleaning is.....
a) b) c) Triethanol ammonium d)
- Shaving soap contains excess of..
a) builders b) perfume c) glycerol d) stearic acids
- Detergents are use ful even in hard water because.....
a) They do not react with hard water ions b) They react with hard water ions but do not form precipitate
c) They destroy the hard water ions ~~Mg~~ d) They sediment undesirable ions in hard water
- The catalyst used in the Hydrogenation of oils is....
a) Al b) Ni c) Zn d) Mg
- Which one is a unsaturated fatty acid
a) Oleic acid b) stearic acid c) Palmitic acid d) None
- The triesters of glycerol are called
a) Fatty acids b) Detergents c) Oils d) Soaps
- Saponification of oils produces
a) Ethanol b) Glycerol c) Fatty acid d) Glycol
- OH groups present in glycerol are
a) 1 b) 2 c) 3 d) 4

KEY

- 1) c 2) b 3) c 4) d 5) b 6) b 7) a 8) c
9) b 10) c

10

UNIT - X

Chemistry and Industry

1. The second largest in terms of population is:

A) China	B) United States	C) Russia	D) India
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1. Cement is mixture of:

a) Sodium silicate and gypsum	b) Calcium silicates and calcium aluminates
c) Sand, clay and felspar	d) Calcium carbonates and sand
2. Glass - blowing is possible with:

a) Flint glass	b) Pyrex glass	c) Soda glass	d) Hard glass
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3. Terra - cotta articles are:

a) Glazed	b) Porous	c) Hard	d) Soft
-----------	-----------	---------	---------
4. Chromophore:

a) soaks the the fibre	b) binds the dye to fibre
c) impart colour to the fibre	
d) intensifies the column of the dye	
5. Drugs which act on blood circulation are:

a) Hormones	b) Vitamins	c) Cardio-vascular	d) antibiotics
-------------	-------------	--------------------	----------------
6. Chief component of cooking gas:

a) butane	b) ethane	c) methane	d) octane
-----------	-----------	------------	-----------
7. Which of the following is a mixed fertilizer?

a)	b)	c) KNO_3	d) Nitrophosk
----	----	-------------------	---------------
8. The type of glass used for making bottle is:

a) Pyrex	b) Boro - silicate	c) Soda glass	d) Quartz
----------	--------------------	---------------	-----------
9. The first synthetic dye prepared by the scientist:

a) W.H.Perkin	b) Fleming	c) Aspidin	d) C.V. Raman
---------------	------------	------------	---------------
10. An example for Auxochrome is:

a) -NO	b) -NO ₂	c) -SO ₃ H	d) c = s
--------	---------------------	-----------------------	----------
11. During setting of cement which of the following acts as binder?

a) CaO	b) CaCO ₃	c) Ca(OH) ₂	d) CaSiO ₃
--------	----------------------	------------------------	-----------------------
12. Glass is a mixture of silicates of:

a) Na, Mg	b) Mg, Ca	c) Na, Zn	d) Na, Ca
-----------	-----------	-----------	-----------

29. Chromophore of Azo dyes:
a) $-\text{NO}_2$ b) $-\text{NO}$ c) $\text{N} = \text{N}$ d) $\text{C} = \text{C}$
30. Dyes having $\text{C} = \text{o}$ as chromophore:
a) Nitro dyes b) Nitroso dyes c) Azo dyes d) Quinone dyes
31. Convert into mordant dyes on reacting with;
a) an acid b) a base c) metal ions d) non - metallic ions
32. Insulin is a:
a) Vitamin b) Chemotherapeutic drug
c) Cardio vascular drug d) Hormone
33. The common binder used in Table making :
a) Gum Arabica b) Epoxy Resin c) Gelatin d) None
34. Acetyl salicylic acid is:
a) Aspirin b) Paracetamol c) Tetracyclin d) Brufen
35. Chemically paracetamol is:
a) o - Hydroxy acetanilide b) p - Hydroxy acetanilide
c) m - Hydroxy acetanilide d) None
36. Gasoline is:
a) Kerosene b) Petrol c) Diesel oil d) Petroleum
37. The by-product in refining petroleum before fractionation?
a) Pitch b) Road Tar c) Gall d) Asphalt
38. Cracking of heavier fraction of petroleum products is for making:
a) Petrol b) Pitch c) Diesel d) Crude oil
39. The main constituent of cooking gas;
a) Methane b) Propane c) Butane d) Ethane
40. Natural gas contains mainly:
a) Methane b) Ethane c) Propane d) Butane
41. Which of the following is a primary nutrient of plants:
a) K b) S c) Mg d) Mn
42. Which of the following is a secondary nutrient of plants?
a) N b) P c) Zn d) Na
43. Which of the following is a micronutrient of plants?
a) K b) S c) Mg d) Fe
44. Microfertilizers contain which of the following?
a) Na b) K c) B d) Mg
45. Commercially available mixed fertilizer:
a) Urea b) Nitrophos K
c) Di and tri ammonium phosphate d) Factumphos

46. Cement is a mixture of silicates and aluminates of:

a) Na

b) Ca

c) Mg

d) Zn

KEY

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|-------|-------|-------|-------|-------|-------|-------|
| 01) B | 02) B | 03) B | 04) C | 05) C | 06) A | 07) D |
| 08) C | 09) A | 10) C | 11) C | 12) D | 13) A | 14) C |
| 15) D | 16) A | 17) C | 18) A | 19) B | 20) A | 21) C |
| 22) B | 23) A | 24) C | 25) B | 26) C | 27) C | 28) B |
| 29) D | 30) D | 31) C | 32) D | 33) C | 34) A | 35) B |
| 36) B | 37) D | 38) A | 39) C | 40) A | 41) A | 42) D |
| 43) B | 44) C | 45) B | 46) B | | | |